



DP IB Environmental Systems & Societies (ESS): SL



Your notes

6.4 Stratospheric Ozone

Contents

- * UV Radiation
- * The Montreal Protocol
- * Stratospheric Ozone



Your notes

UV Radiation

UV Radiation Effects

- The Sun emits electromagnetic radiation in a range of wavelengths, from low-frequency radio waves to high-frequency gamma radiation
- **Shorter wavelengths** of radiation have higher frequencies
 - **More energy damages living organisms**
 - E.g. ultraviolet (UV) radiation

Effects on human health

- Ultraviolet radiation from the Sun can have damaging effects on **human living tissues**
 - When excessive UV radiation reaches the surface of the Earth, it can lead to various health issues by damaging cells and tissues

UV Radiation Effects on Humans

Health issues caused by UV radiation	Explanation
Cataracts	<p>Prolonged exposure to UV radiation can contribute to the development of cataracts</p> <p>Cataracts cause clouding of the lens in the eye, leading to blurry vision and eventual vision loss if left untreated</p>
UV radiation affects cells	<p>UV radiation has the potential to induce mutations in DNA during cell division</p> <p>When cells are exposed to UV radiation, it can lead to genetic alterations and mutations</p> <p>This can disrupt normal cell growth and increase the risk of developing cancer</p>
Skin cancer	<p>UV radiation is a major risk factor for the development of skin cancer</p> <p>UV rays can damage the DNA in skin cells, leading to uncontrolled cell growth and the formation of cancerous tumours</p>



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	Prolonged or intense exposure to UV radiation, especially without proper protection , increases the risk of developing skin cancer
Sunburn	When the skin is exposed to excessive UV rays, it triggers an inflammatory response as a defence mechanism Sunburned skin becomes red, painful and may blister, indicating damage to the skin cells
Premature skin ageing	Chronic exposure to UV radiation accelerates the ageing process of the skin It can cause the breakdown of collagen and elastin fibres, leading to wrinkles , sagging skin and the development of age spots



Measures such as wearing sunscreen with a high sun protection factor (SPF), using sunglasses that block certain UV rays, seeking shade during peak sun hours and wearing protective clothing can help reduce the risk of UV-related health issues (Photo by Dimitris Chapsoulas on Unsplash)

Effects on biological productivity

- Harmful UV radiation reaching the Earth's surface affects plant growth and productivity
- Increased UV exposure can lead to:
 - Reduced photosynthesis rates
 - Altered plant metabolism
 - Decreased crop yields
- Exposure to increased UV radiation can affect other photosynthetic organisms, such as **phytoplankton**
 - Phytoplankton play a crucial role in **aquatic food webs**
 - They convert sunlight, carbon dioxide and nutrients into organic matter through **photosynthesis**
 - UV radiation damages phytoplankton by:
 - Causing DNA damage
 - Reducing photosynthetic activity and growth
 - This leads to a decrease in primary productivity in aquatic ecosystems
- Reduced phytoplankton productivity can have cascading effects on **higher trophic levels** in aquatic ecosystems
 - **Zooplankton**, which feed on phytoplankton, have less food available
 - This affects their growth and reproduction
 - This, in turn, can impact higher-level consumers, such as **fish** and **marine mammals**
 - Organisms in these higher trophic levels rely on phytoplankton and zooplankton as food source
 - This can significantly reduce the biodiversity of aquatic ecosystems



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The Montreal Protocol



Your notes

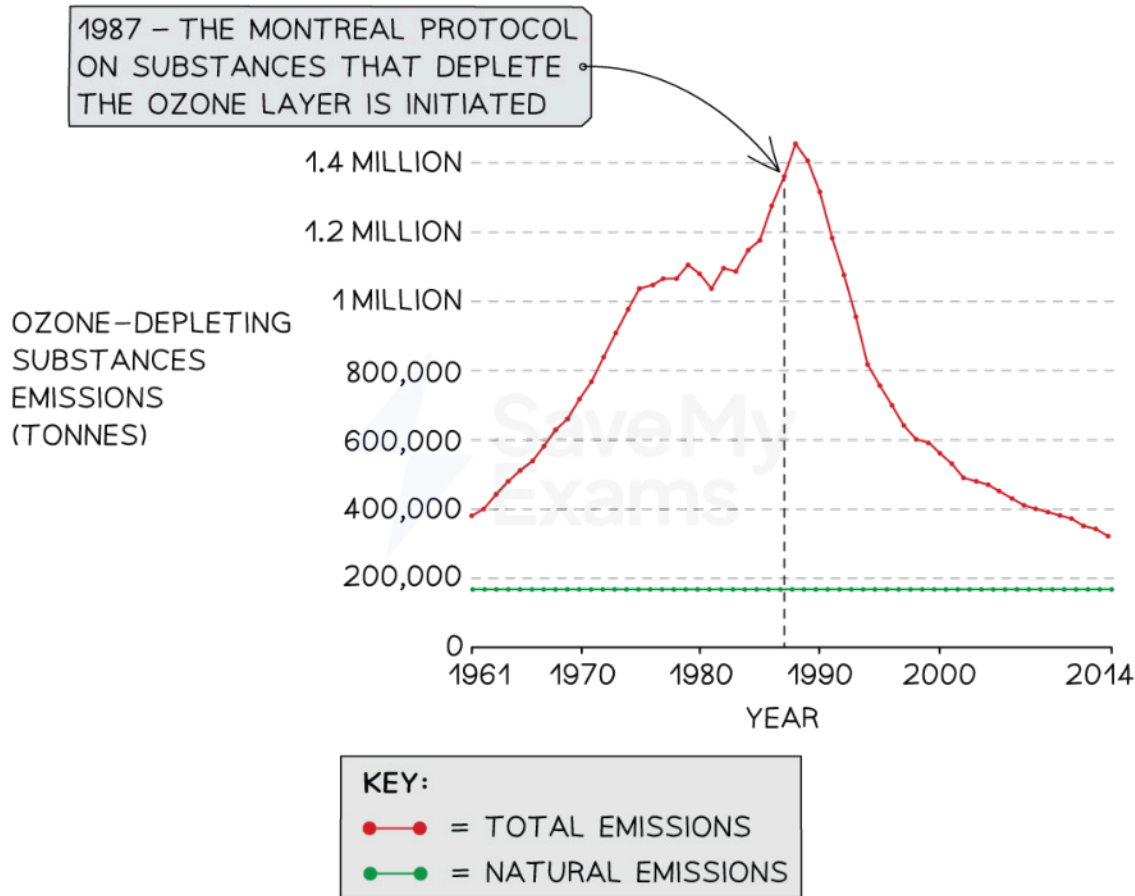
The Montreal Protocol

The role of UNEP

- The United Nations Environment Programme (UNEP) has played a critical role in the **protection** of the **stratospheric ozone layer**
 - This have been achieved through its efforts in providing information and creating **international agreements**:
- UNEP has been instrumental in raising awareness about:
 - The fact that the ozone layer was being rapidly depleted
 - The causes of this depletion
 - The associated environmental and health impacts of this depletion
 - Through research and sharing of information, UNEP has helped educate governments, industries and the public about the importance of **ozone layer protection**
 - UNEP has been actively involved in the creation of international agreements aimed at **reducing** the use of **ozone-depleting substances** (ODSs)



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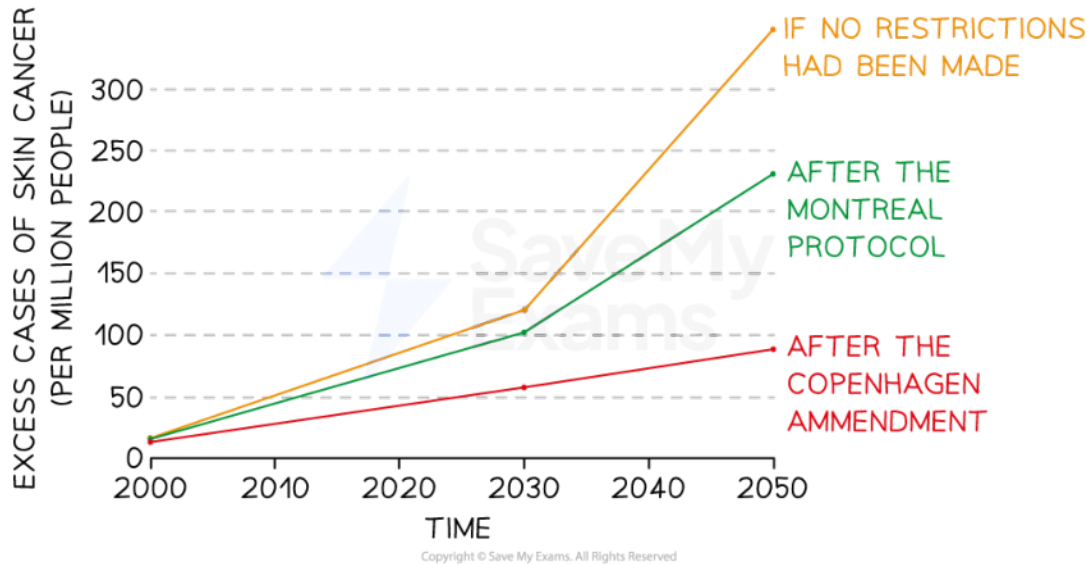
The Montreal Protocol on Substances that Deplete the Ozone Layer, initiated in 1987, played a crucial role in reducing the production, use and emissions of ozone-depleting substances (natural sources of halogen gases that deplete stratospheric ozone include biological activity in terrestrial and aquatic ecosystems)

- The **Montreal Protocol on Substances that Deplete the Ozone Layer** was initiated in 1987
 - It was started under the guidance of UNEP
 - It is a landmark international agreement that regulates the production, trade and use of chlorofluorocarbons (CFCs) and other ODSs.
 - 24 countries initially signed the initial protocol, and the total now stands at **197 countries**
 - It has been updated and strengthened (a later amendment at a summit in Copenhagen in 1992 tightened restrictions further)



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- It has resulted in emissions of ODSs **falling rapidly** from around 1.5 million tonnes in 1987 to around 400 000 tonnes in 2010
- UNEP hopes to end production of all HCFCs by 2040



The effects of CFC reduction on predicted skin cancer rates in Northwest Europe

- The **illegal market** for ozone-depleting substances is a significant challenge to the effectiveness of ozone protection efforts:
 - UNEP recognises the need for consistent **monitoring** and **enforcement** to tackle this issue
 - By collaborating with national authorities, customs agencies and other relevant stakeholders, UNEP works towards:
 - Stopping the illegal trade of ozone-depleting substances
 - Ensuring **compliance with international regulations**
- Phased reductions:**
 - Gradual reduction schedules for ODSs have allowed industries to adapt
 - The Montreal Protocol provided time for the development and adoption of alternatives to ODSs
- National governments** play an important role in implementing the agreements made by the UNEP:
 - In response to the Montreal Protocol, governments have enacted **national laws** and **regulations** to decrease the consumption and production of halogenated organic gases, such as chlorofluorocarbons (CFCs)

- These laws help enforce the reduction targets and promote the **transition to ozone-friendly alternatives**
- The collective efforts of UNEP, governments, industries and other stakeholders are vital in achieving goals, including:
 - Ozone layer protection
 - Mitigating the illegal trade of ozone-depleting substances
 - Encouraging global cooperation for a more sustainable future



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Planetary boundary for stratospheric ozone depletion

- Stratospheric ozone depletion is one of the nine planetary boundaries outlined by the [planetary boundaries model](#)
 - Planetary boundaries are **thresholds** that lead to significant environmental changes if they are crossed
- The Montreal Protocol is regarded as the most successful example yet of international cooperation in management and intervention to resolve a significant environmental issue
 - Actions taken in response to the Montreal Protocol have prevented the planetary boundary for stratospheric ozone depletion being crossed
- **Evidence from data:**
 - Data shows a decrease in the size of ozone holes over time
 - Continuous monitoring indicates that ozone layer recovery is underway

EXAM TIP



Make sure you understand the key reasons for the Montreal Protocol's success (e.g. international cooperation, legally binding agreements, etc.) and that you are familiar with how the protocol has helped reverse stratospheric ozone depletion.



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Stratospheric Ozone

Stratospheric Ozone

- Ozone is a molecule composed of three oxygen atoms (O_3)
 - It is mainly found in the Earth's **stratosphere**
 - This is a layer of the atmosphere located approximately 10 to 50 kilometres above the Earth's surface
- Ozone plays a very important role in **protecting life** on Earth
 - This is because it **absorbs** a significant portion of the Sun's harmful UV radiation
 - This significantly reduces the amount of UV radiation that reaches the Earth's surface
- **Types of UV radiation:**
 - **UVA:**
 - Longest wavelength
 - Least harmful but can cause skin aging and contribute to skin cancer
 - **UVB:**
 - Medium wavelength
 - Can cause skin burns and direct DNA damage
 - Mostly absorbed by stratospheric ozone, but some reaches the Earth's surface
 - **UVC:**
 - Shortest wavelength
 - Most harmful
 - Completely absorbed by stratospheric ozone

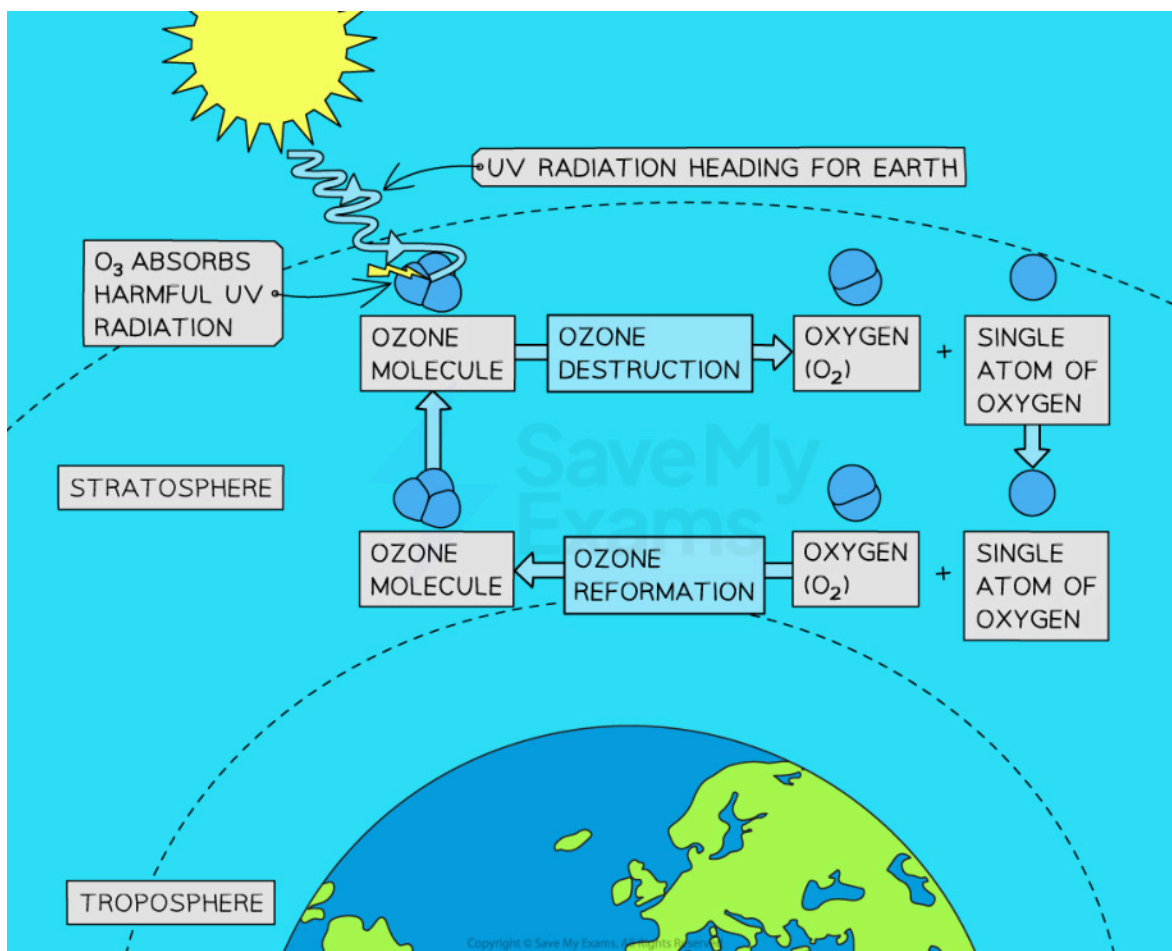
Ozone Equilibrium

- The amount of ozone in the stratosphere remains relatively constant over long periods
 - This is due to a **steady state** of **equilibrium**
- Equilibrium is maintained between the processes of ozone **formation** and **destruction**



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- When UV radiation from the Sun interacts with ozone molecules, some of the ozone **absorbs** the energy and **breaks apart**
- This results in the formation of an oxygen molecule (O_2) and a free oxygen atom (O)
 - This process of ozone destruction occurs **naturally** in the stratosphere
 - Under normal conditions, the free oxygen atom (O) can combine with another oxygen molecule (O_2) to form ozone (O_3) again
 - This ozone destruction and reformation creates a **dynamic equilibrium** in the stratosphere
 - There is a **continuous cycle** of ozone molecules being broken apart and reformed
 - This dynamic equilibrium ensures that the concentration of ozone in the stratosphere remains relatively **stable** over time
 - The rate of the forward reaction equals the rate of the backward reaction in the system, so the concentrations of the **reactants** and **products** remain relatively **constant**



Ozone destruction and reformation creates a dynamic equilibrium in the stratosphere—there is a continuous cycle of ozone molecules being broken apart and reformed



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Ozone Depletion

Ozone-depleting substances

- Ozone-depleting substances (ODSs) are chemicals that cause **stratospheric ozone depletion**
 - These substances cause the destruction of ozone molecules
 - In other words, they enhance the natural ozone breakdown process (beyond natural levels)
- ODSs are commonly used in various human activities and products:

Sources of Ozone Depleting Substances

Source	Details
Aerosols	<p>Chlorofluorocarbons (CFCs) were previously used as propellants in aerosol products like sprays, foams, and deodorants</p> <p>When released into the atmosphere during spraying, these substances can eventually reach the stratosphere and contribute to ozone depletion</p>
Gas-blown plastics	<p>ODSs were also used as blowing agents in the production of foamed plastics</p> <p>These agents help create air pockets within the plastic material, making it lightweight</p> <p>During manufacturing or disposal of these products, ODSs can be released into the atmosphere</p>
Pesticides	<p>Some pesticides, e.g. those containing methyl bromide, have been used in agricultural practices for soil fumigation</p> <p>When applied, these substances can vaporise and enter the atmosphere, where they can contribute to ozone depletion</p>
Flame retardants	<p>Some flame retardants contain halogen atoms and have been used in various products to reduce their flammability</p> <p>When these products degrade or are disposed of, the halogenated compounds can be released into the atmosphere</p>



Your notes

Refrigerants	<p>ODSs were widely used as refrigerants in cooling systems, such as air conditioners and refrigerators</p> <p>The most well-known examples are CFCs</p> <p>When these refrigerants leak or are improperly disposed of, they can reach the stratosphere and contribute to ozone depletion</p>
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Imbalance in equilibrium

- When ozone formation and destruction rates are **unequal**, equilibrium is **disrupted**
 - This leads to increased ozone depletion
 - Increased **UVB** radiation reaches the Earth's surface
 - Affects **ecosystems** and **human health**
 - Causes increased rates of skin cancer and cataracts
 - Reduces terrestrial and marine productivity

Ozone holes

- Ozone depletion affects the entire Earth's stratosphere
 - However, **ozone holes** are most prominent at the poles
 - Ozone holes are areas of **low stratospheric ozone**
 - These holes appear every spring due to ODSs and seasonal weather patterns

EXAM TIP



The use of chemical symbols, formulae or equations for the equilibrium of ozone is not required, so you do not need to learn the chemical equations relating to the formation and destruction of ozone.